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54 **In situ crosslinking of polyelectrolytes.**

57 A composition, useful for forming absorbent articles of a carboxylic polyelectrolyte, is prepared by copolymerizing a carboxylic polyelectrolyte monomer, a reactive self-cross-linking monomer and a hydrophobic monomer. The composition in the form of an emulsion or a viscous solution has an extendable shelf life and may be used to produce a superabsorbent material or a substrate impregnated with a superabsorbent material.

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IN SITU CROSSLINKING OF POLYELECTROLYTESBackground of the Invention

5 This invention relates to products for absorbing
aqueous fluids and, in particular, products for absorbing
body fluids. The invention is specifically directed
toward compositions for producing crosslinked polyelectro-
lytes which compositions are usable in commercial products
10 for absorbing aqueous liquids and are particularly useful
in producing absorbent composite materials which are
coated or impregnated with such compositions and
subsequently treated so as to produce a composite material
incorporating polyelectrolyte which is crosslinked in situ.

15 Highly absorbent crosslinked polyelectrolytes and
methods for preparing the same are already known. U.S.
Patent Nos. 3,669,103 and 3,670,731 teach the use of these
materials in diapers and dressings. U.S. Patent Nos.
20 2,988,539; 3,393,168; 3,514,419 and 3,557,067 teach
methods of making such absorbents and, in particular, are
related to making water swellable crosslinked carboxylic
copolymers that are either crosslinked during copolymeri-
zation or crosslinked after polymerization and then
25 neutralized to result in pendant ionic moieties capable of
imparting water retention properties to a finished
material.

 Further, the process of incorporating such poly-
30 electrolytes or their monomeric precursors into a
substrate such as a fibrous web or a cellular, i.e. foam,
material and then crosslinking in situ, is also known.
Published European Patent Office Application 81302086.4
discloses impregnating a fibrous web with a solution of
35 monomers, with or without crosslinking agents and then

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irradiating such impregnated web to polymerizing and/or crosslink in situ. U.S. Patents 4,076,673; 3,980,663; 4,071,650 all exemplify teachings of preparing solutions for coating substrates which subsequently may be treated to produce in situ crosslinked polyelectrolytes.

Unfortunately, these foregoing suggestions suffer from several drawbacks when it is contemplated to translate such suggestions into commercial processes. Suggestions directed toward irradiation techniques, when scaled to commercial proportions, require large capital investment for irradiating equipment such as, for example, an electron beam device, and, concomitantly, consume large quantities of energy for carrying out such processes. Additionally, in order to obtain high conversion of monomer to polymer, such irradiation technique requires still more energy input and also tend to over crosslink the polymer resulting in a decrease in absorption capacity.

Those suggestions relating to preparing compositions comprising polyelectrolytes and crosslinking agents and then applying such composition to a substrate are also fraught with problems on a commercial level. A major drawback to such systems is the limited shelf life of the impregnating or coating concentration. Almost immediately upon mixing the polyelectrolyte with the crosslinker, a reaction begins that leads to gel formation. Once gelled the composition can no longer be used. Accordingly, a commercial process must, to utilize such suggestions, apply the components i.e., the polyelectrolyte and the crosslinker, in separate steps and hence suffer the concomitant inconvenience as well as the likelihood of locally applied non-uniform proportions of these components. The problem is compounded when such commercial processes are designed to impregnate a substrate

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with excess composition and then recirculate to reuse the excess. Such a system will be encountered when attempts are made to impregnate a fibrous or cellular web with vacuum deposition techniques such as are common in the art of producing bonded non-woven fabrics. In such a process, of necessity, the active components are intermixed and gelling follows.

Accordingly, there is a need for providing a commercial process for in situ crosslinking of polyelectrolytes.

Summary of the Invention

In accordance with the teachings of this invention a composition and process is provided that can be employed for coating or impregnating a substrate, which will deposit polyelectrolyte onto or into the substrate and which polyelectrolyte can be crosslinked by subsequent drying or heating in situ. The composition is so selected as to have a long shelf life and will not gel prior to use or after it is collected as a reusable, recirculated composition.

Specifically, the invention contemplates utilizing, with modification, a technique employed in the art of manufacturing acrylic fibers such as are produced from spinnable solutions of copolymers comprising acrylonitrile. In U.S. Patent Nos. 4,059,556 and 4,100,143; for example, it is taught that non-poly(electrolytic) acrylonitrile copolymers may be provided which are self crosslinking by the inclusion therein of a comonomer which is capable of crosslinking upon drying or heating. Such comonomers suggested are exemplified by N-methyl alkyl ether acrylamides and N-methylol compounds or N-methyl ethers of unsaturated mono- or bis-urethanes. As is taught in the art of fiber production exemplified by these

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patents, emulsions of such self crosslinking copolymers are employed and may be stabilized by either incorporating additives into the emulsion e.g. mineral salts, α -amino carboxylic acids or by effecting the copolymerization reaction in strongly polar organic solvents such as dimethyl formamide or dimethyl sulphoxide. It is also mentioned in these references that additional optional monomers may be copolymerized into the polymer, such optional monomer having been conventionally employed in the acrylonitrile fiber technology and exemplified by such monomers as acrylic acid and methacrylic acid alkyl esters, among others. The purpose of these monomers, as has been employed in the acrylonitrile fiber technology is to render the copolymer crosslinkable.

15

It has now been discovered that a composition, useful for forming absorbent articles comprising crosslinked carboxylic poly(electrolyte) may be provided by employing an emulsion of the poly(electrolyte) copolymerized with monomers of the kind that produce self crosslinking copolymers, as employed in the acrylic fiber art. Moreover, it has been discovered that the problem of emulsion instability is cured in the case of self crosslinking poly(electrolyte) copolymers without resorting to undesirable additions or to employing, for the dispersing phase of the emulsion, organic solvents.

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Indeed, it has been discovered that when attempting to utilize the teachings of the acrylic fibrous art, in the use of polyelectrolytes, a self crosslinkable copolymer cannot be conveniently prepared. The presence of the acidic groups in the electrolyte monomers tends bringing about premature gelling, i.e. crosslinking, during the copolymerization reaction and renders the composition useless for the purposes of this invention.

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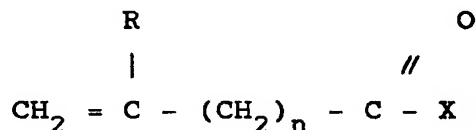
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It has now been discovered that the techniques of producing self crosslinkable copolymers may be applied to the art of in situ crosslinking of polyelectrolytes, provided that, in addition to the polyelectrolytes and reactive self crosslinking monomers, the copolymer also includes an effective proportion of hydrophobic comonomers.

Specifically, it has been discovered that a composition, useful for forming absorbent articles of a carboxylic polyelectrolyte may be prepared which composition comprise a polymer which is copolymerized from:

- (a) a carboxylic polyelectrolyte monomer;
- (b) a reactive self crosslinking monomer having the formula:



wherein R is selected from the groups consisting of H or CH₃, n is selected from the group consisting of zero or the integers 1 to 5, and X is selected to be a group capable of reacting through self condensation or with carboxylic acid by condensation or ring opening; and

- (c) a hydrophobic monomer.

As used herein, hydrophobic monomer is meant to define a water immiscible monomer which, when present in sufficient quantity, will maintain the un-neutralized acidic polymer resulting from the above composition in the emulsion state. The hydrophobic monomer preferably is selected as an alkyl ester of acrylic or methacrylic acid, said alkyl group having from one to eight carbon atoms, and preferably from one to four carbon atoms. Preferably, the

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reactive self crosslinking monomer, (b), is one in which X is selected from the group consisting of glycidyl, N-methylol, N-methylol alkyl ether having alkyl groups of from one to four carbon atoms, blocked isocyanate, 5 N-methylol urethane, or aziridine.

In many cases, depending upon the reactivity of the crosslinkable monomer selected, the emulsion resulting from the copolymerization reaction, of the above-described 10 composition is stable and will not crosslink after a substantial period of time. In some cases, however, it is preferred, soon after forming such emulsion, to neutralize the composition to a pH of greater than 5 by employing, for example, an aqueous sodium hydroxide solution. This 15 neutralization renders the copolymer soluble, breaks the emulsion, and has been found to substantially increase the shelf life of the resulting solution without affecting its ability to self crosslink upon drying and/or heating.

20 The solution is employed by coating or impregnating various substrates and then drying and/or heating to result in in situ crosslinked absorbent polymer.

Detailed Description of the Invention

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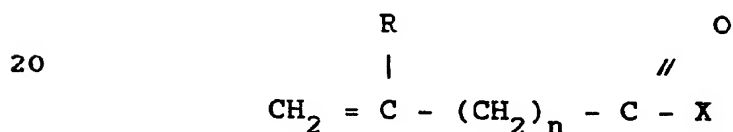
The invention concerns self crosslinking polyelectrolyte copolymers of polyelectrolyte moieties, self crosslinking reactive moieties and hydrophobic moieties.

30 The carboxylic polyelectrolyte moieties capable of being crosslinked in accordance with the teachings of this invention are well known and are described in detail in U.S. Patent No. 4,310,593 which is incorporated herein by reference. The essence of usable polyelectrolytes is that 35 they comprise, at least in the salt form, sufficient

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carboxylate groups to render them water soluble. Usable polymers, capable of being prepared from readily available monomers and converted into their salt form include, for example, acrylic acid-acrylate copolymers; acrylic acid-acrylamide copolymers; acrylic acid-olefin copolymers; poly acrylic acid; acrylic acid-vinyl aromatic copolymers; acrylic acid-styrene sulfonic acid copolymers; acrylic acid-vinyl ether copolymers; acrylic acid vinyl acetate copolymers; acrylic acid-vinyl alcohol copolymers; copolymers of methacrylic acid with all of the above comonomers; copolymers of maleic acid, fumaric acid and their esters with all of the above comonomers; copolymers of maleic anhydride with all of the above comonomers.

The self crosslinking reactive moieties are provided by copolymerizing together with the polyelectrolyte, reactive groups having the formula:



wherein R is selected from the group consisting of H or CH_3 , n is selected from the group consisting of zero or the integers 1 to 5, and X is chosen to be a group capable of reacting either through self condensation or, with carboxylic acid, through condensation or a ring opening mechanism.

Thus, for example, X may be selected from the group consisting of glycidyl; N-methylol; N-methylol alkyl ether wherein the alkyl group has from 1 to 4 carbon atoms; blocked isocyanates such as, for example, an isocyanatoethyl methacrylate - caprolactam adduct, N-methylol urethane; aziridine or the like.

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Reactive groups such as are described above have already been incorporated in copolymers to produce self crosslinking polymers (see for example U.S. 4,100,143 and U.S. 4,059,556). However, such self crosslinking polymers as have employed these groups are primarily hydrophobic in nature. The present invention, instead, involves self crosslinking copolymers wherein the carboxylic groups (acids and salts thereof) represent a substantial portion of the copolymer and specifically range from 20% to 82% of the copolymer, on a weight basis and preferably from 35 to 75 percent, by weight.

The presence of this high proportion of carboxylic groups, while necessary to result in a polyelectrolyte having a high capacity for absorbing aqueous fluids, has a concomitant disadvantage, when copolymerized with the reactive, self crosslinking groups as set out above. It has been discovered that almost immediately after the copolymer reaction is carried out, because of the presence of this high proportion of carboxylic groups, the crosslinking reaction, being catalyzed by acid, begins. The result is that the copolymerized composition quickly turns to gel. For example, a copolymerization reaction of meth(acrylic) acids with 1 to 2%, by weight, glycidyl methacrylate or with N-methylol acrylamide, carried out in water at 70°C, turned into a gel in about twenty minutes. Clearly such composition is not conveniently usable for the purposes described herein.

In accordance with the teachings herein, it has been discovered that this problem may be greatly ameliorated by copolymerizing a third moiety group into the self crosslinking polymer. Generally, said third moiety is a hydrophobic monomer i.e., a water immiscible monomer. The hydrophobic monomer is copolymerized into the self

crosslinking polyelectrolyte chain in sufficient quantity to preclude premature gelation. While such quantity may vary with such features as the specific moieties chosen for all three types of copolymers, generally such quantity
5 should represent about 5 to about 70% by weight of the finished poly(electrolyte) self crosslinking copolymer. Preferably such range should be from about 10 to about 35%. Suitable hydrophobic comonomer may be selected from the group consisting of the alkyl esters of acrylic and
10 methacrylic acid wherein said alkyl groups consist of from one to four carbon atoms.

It has been discovered that by incorporating such above-described hydrophobic groups into the polyelectro-
15 lyte self crosslinking copolymer, the resulting composition has a significantly greater shelf life wherein the tendency to prematurely crosslink and gel is greatly reduced. Without being bound to any particular theory, it is believed that the primary function of these hydrophobic
20 moieties is to maintain the reaction product of the copolymerization process in the dispersed phase of an aqueous continuous phase emulsion when the reaction is completed. In the emulsion form, the copolymers produced are coiled, with a preponderance of the crosslinkable
25 groups hidden inside the colloidal particles and, hence, less accessible for the crosslinking reaction.

The retardation of the crosslinking reaction is effective in all cases by incorporating the hydrophobic
30 moiety. Still greater retardation may be achieved by employing the teachings of this invention with further modification. Since crosslinking is retarded, there is now sufficient time to complete the copolymerization reaction and then further treat the reaction mixture by
35 neutralizing the acid functions to a pH of about 5 or

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greater. The emulsion is broken by such neutralization and a clear viscous solution results. This neutralized solution has been found to have exceptionally long shelf life and is preferred when such long life is a consideration. Needless to say, it is important to first copolymerize the copolymers of polyelectrolyte and self crosslinking moieties with the hydrophobic polymeric moieties, as taught above, in order to have sufficient time to carry out the subsequent neutralization reaction without encountering premature gelation. The neutralization step is most preferably used when employing as the self crosslinking moieties, a particularly reactive type, such as for example a glycidyl moiety.

Preferably, the reaction is carried out in water. The reaction may be carried out at pressures ranging from subatmospheric to super atmospheric pressure and at temperatures of from 15°C to reflux, and may be catalyzed by such free radical generating compounds such as peroxide or azo compounds. Preferably the reaction is accomplished by using an oxidation/reduction initiator such as, for example, persulfate, bisulfate and ferrous ions. When this latter method is employed, a relatively low polymerization temperature will be sufficient to carry out the reaction. Low temperatures will also aid in the retardation of premature gelation.

The composition of this invention is useful for producing a wide variety of product and material having the common property of absorbing large quantities of aqueous liquids. For example, one employment of the composition is to produce a powdery "superabsorbent", i.e. an absorbent material capable of absorbing at least ten times its own weight of aqueous liquid. Such powder may be obtained by the mere heating and/or drying of the

composition, prepared in accordance with the teachings of this invention either in the form of an emulsion or neutralized solution. The composition is cast upon a substrate to form a crosslinked polyelectrolyte film which is subsequently crushed into a powder. It should be noted, of course, that the film itself has substantial absorbent properties and is useful in some applications directly without the need for crushing into a powder.

The composition of this invention, when employed in the manner indicated above to produce a powder or film, manifests certain unique advantages. Typically, prior synthesis of superabsorbent materials involved the steps of synthesizing the polymer in question in an aqueous medium and then drying the reaction mixture to provide a dry superabsorbent powder. Such prior process are described, for example, in U.S. Patent Nos. 4,090,013 and 4,320,040. The drawback in such a system is that the superabsorbent is formed in the presence of large quantities of water and, owing to the highly absorbent properties of the polymer, a great quantity of water will be tightly bound to the polymer. Accordingly, drying to a powder, for example, is exceedingly difficult and involves a high energy input. On the other hand, the obvious alternative is to use an organic solvent. Unfortunately the drawback is high cost, high volatility and the danger of explosion.

In contrast to these prior methods, by employing the composition of this invention, the initial removal of water by heating or otherwise, simply concentrates the polymer and does not effect crosslinking. Instead, it is only in the relatively dry state, when most of the water has been removed, that crosslinking occurs and concomitantly, the polymer becomes a "super absorbent". Since

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this only occurs in the relatively dry state, the quantity of tightly bound water to be removed is substantially less than in prior processes for manufacturing superabsorbents and, hence, far less difficult to accomplish.

5

In another embodiment of this invention, the composition made in accordance with the teachings herein is employed to enhance the aqueous liquid absorption characteristics of an aqueous liquid impregnable substrates. Such substrates may comprise, for example, fibrous webs of cellulosic or synthetic fiber, either woven or non-woven, or loosely associated batts of such fibrous material. Additionally, such substrates may comprise cellular polymeric material, i.e. foams such as for example, polyurethane foam or regenerated cellulose foam. In such embodiment the composition of this invention is sprayed, padded, flooded or otherwise applied to the substrate and preferably impregnated into the substrate by means well known in the art such as applying a differential pressure by mechanical compression or vacuum application. The application of the composition to the web may be selective, i.e. not uniform over the entire surface of the web but instead in a pattern such as a grid pattern, elongated stripes of a desired width, spots, or the like. Additionally, the depth of penetration into the Z-direction of the web may be controlled by, for example, appropriate control of the vacuum applied, where vacuum deposition is the selected deposition method. The impregnated web is then dried and/or heated, and the deposited copolymer crosslinks, in situ, to provide a substrate of enhanced absorbency.

It will be apparent to one skilled in the art that the production of substrates incorporating deposited superabsorbent material are made commercially in large volume.

generally by treating a continuously moving web of substrate material. It will also then be apparent that the unique long shelf life of the composition of this invention is particularly advantageous to such
5 production. Further, the composition of this invention will greatly facilitate the drying of the impregnated substrate. Prior methods involving in situ polymerizing and/or crosslinking of deposited material on a substrate have been based on processes whereby the polymerization
10 and/or crosslinking occurs while substantial quantities of the dispersing phase, generally water, is still present. Accordingly the final polymer is a superabsorbent material which will, because of the presence of large quantities of water, absorb and retain such water aggressively. There-
15 fore, drying of the substrate has also been difficult. In contrast therewith, by employing the composition of this invention, the self crosslinking does not actually occur until a great preponderance of the water present in the composition has been first removed. It is only at this
20 point that the copolymer becomes a superabsorbent and, hence, less of a quantity of tightly held water need be removed by following the teachings of this invention.

The products described above, e.g. the super absorbent
25 powder, film and substrates are useful in the manufacture of various products having in common the property of employment for absorbing liquids in general, and for absorbing body fluids in particular. Typically, the products of this invention may be used as the absorbent
30 element or as one of several absorbent elements in products such as sanitary napkins, catamenial tampons, incontinence pads, underpads for beds and the like, dressings, clothing shields such as panty liners or underarm dress shields, wound dressings, disposable
35 diapers and similar products.

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In order to better understand the teachings of this invention, the following examples are given:

Comparative Example 1

5 Seven hundred and fifty mls. of distilled water are charged into a one liter flask which is equipped with agitation equipment, a nitrogen purge source and a thermocouple. The liquid is de-aired for a half hour with
10 nitrogen to dispel any dissolved oxygen which might otherwise inhibit the copolymerization reaction. During this time, the water is heated to 70°C. After reaching this temperature, 0.5 g of potassium/sulfate are dissolved in 20 mls. of distilled water and added to the reaction
15 vessel. Within one minute 80 mls. of acrylic acid, having a purity of 99% by weight and 20 mls. of methacrylic acid having a minimum purity of 98%, by weight, are added to the reaction vessel. Immediately following the addition of these monomers, 1 ml. of glycidyl methacrylate, the
20 reactive crosslinkable monomer, is added to the reaction vessel. The reaction is allowed to proceed for 30 minutes. At this point, it is noted that the reaction mixture has crosslinked and formed an unworkable gelled material.

25

Comparative Example 2

30 The procedure of Example 1 is repeated with the exception that in lieu of the 1 ml. of glycidyl methacrylate as the reactive crosslinkable monomer, 2 mls. of a 60% aqueous solution of N-methylol acrylamide is employed. Again, after heating the reactants for 30 minutes, it is noted that the reaction mixture has crosslinked and an
35 unworkable gelled material results.

- 15 -

As can be seen from the above examples, the copolymer obtained and comprising polyelectrolyte in combination with self crosslinking reactive monomers is not usable in the processes of this invention because of premature
5 gelling.

Example 1

Seven hundred and fifty mls. of distilled water are
10 charged into a one liter flask equipped with agitation means and nitrogen purge means and thermocouple. The liquid is de-aired for one half hour with nitrogen while simultaneously being heated to 70°C. After reaching the desired temperature, the nitrogen purge is removed and
15 positioned over the surface of the solution to maintain a nitrogen blanket. Thereafter, 15 gms. of a 30% by weight aqueous solution of sodium laurel sulfate, a surfactant, are added to the solution together with 0.5 gm. of potassium/sulfate dissolved in a solution of 20 mls. of
20 distilled water. Within one minute, 85 mls. of acrylic acid and 15 mls. of the hydrophobic monomer methyl methacrylate, having a minimum purity of 98% by weight are added to the reaction vessel. Following the addition of these monomers, 1 ml. of glycidyl methyl methacrylate is
25 added to the reaction vessel and the reaction is allowed to proceed for 30 minutes. The reaction mixture resulting from this procedure is a stable, fluid, milky white emulsion. Half of this emulsion is allowed to stand at room temperature, and it is only after two days that it is
30 noted that a small portion of the emulsion has undergone crosslinking and gelled. The second half of the emulsion obtained from the reaction is treated with a 50% by weight aqueous sodium hydroxide solution in sufficient quantity to raise the pH to about 5. Upon the addition of the
35 sodium hydroxide solution a viscous clear solution

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results. This neutralized solution is also allowed to stand at room temperature, and after two days appears to be fluid, workable and non-gelled.

5 The neutralized solution is dried for 15 minutes at 120°C. and a film is obtained. The film is crushed to result in superabsorbent powder. The powder is tested for absorbency using the free absorbency method which comprises placing 0.1 grams of the powder in a beaker
10 containing 15 mls. of a 1% by weight sodium chloride aqueous solution. The dispersed superabsorbent in the sodium chloride solution is filtered, excess water is allowed to drain by gravity, and the remaining material is weighed. It is determined that the powder has absorbed 40
15 gms of the solution per gram of dried powder.

Example 2

20 Again, as in Example 1, 750 mls. of distilled water are charged to the 1 liter flask, the liquid is de-aired with nitrogen and heated to 70°C. The surfactant and potassium per sulfate are added and then, within one minute, 70 mls. of acrylic acid, 20 mls. of methacrylic acid and, as the hydrophobic monomer, 10 mls. of N-butyl
25 acrylate are added. Immediately following the addition of these monomers, 2 mls. of a 60% aqueous N-methylol acrylamide, a less reactive crosslinking monomer than the glycidol methacrylate recited in Example 1, is added. The reactants are heated for 30 minutes and a stable, fluid,
30 milky white emulsion is obtained. Half the emulsion is neutralized with enough 50% by weight sodium hydroxide aqueous solution to result in a clear solution having a pH of 5. The other half remains in the emulsion form. Both portions are stored for one month at room temperature, and
35 it is observed that in this case, employing the less

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reactive self crosslinking monomer, both the emulsion form and the neutralized form are still fluid, workable polymer solutions. The solutions are capable of being employed in each of the processes set out above to produce highly
5 absorbent material.

Example 3

The neutralized solution obtained from the procedures
10 of Example 1 is utilized to construct an absorbent structure by applying it to a polyester fibrous web. A carded polyester fibrous web having fibers of a denier of 8.5, staple fiber length of 1.5 inches, and a basis weight of 1 oz/yd², is saturated with the solution and excess
15 is removed by vacuum suction. The web is then dried for 20 minutes at 120°C. resulting in a structure consisting of 40% by weight of the polyester fibers and 60% by weight of in situ crosslinked polymer. A sample of this material is tested for maximum capacity and fluid retention utiliz-
20 ing a mechanized form of the Porous Plate Testing apparatus as described in detail in Textile Research Journal 37, pp. 356-366, 1967. Briefly, this test involves placing the sample in what is essentially a Buchner funnel having a porous bottom plate and holding the sample in place by
25 applying thereon a standard weight to maintain a standardized confining pressure. In the case of this test, the confining pressure was 0.07 psig. The porous plate is placed in contact with a reservoir of fluid which again for this test comprises a 1% by weight sodium chloride
30 aqueous solution, and the sample is allowed to absorb the fluid through the porous plate until saturated. By maintaining the sample at 1.5 cm above the level of the reservoir to avoid flooding, the fluid absorbed is subjected to essentially zero hydrostatic head with respect
35 to the reservoir. The volume of fluid absorbed, divided

by the weight of the sample, is termed the maximum capacity. To determine fluid retention, the saturated sample is elevated with respect to the reservoir thereby imposing a hydrostatic head upon the fluid absorbed, the head arbitrarily being chosen as 35 cm of fluid. The apparatus is provided with means for measuring the volume of fluid retained under the hydraulic head. Retention values are reported as the volume retained per unit weight of the sample. The results of this test for the web of this example resulted in a maximum capacity of 18 ml/gram of 1% aqueous sodium chloride and a retention value of 11 ml/gram.

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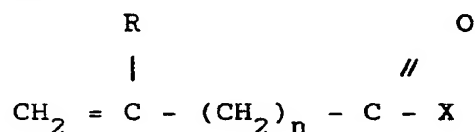
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What is Claimed:

1. A composition, useful for forming absorbent articles employing a carboxylic polyelectrolyte, comprising a poly(electrolyte) self crosslinking copolymer
 5 copolymerized from:

(a) a carboxylic polyelectrolyte monomer;

10 (b) a reactive self crosslinking monomer having the formula:



15

wherein R is selected from the groups consisting of H or CH₃, n is selected from the group consisting of zero or the integers 1 to 5, and X is selected to be a group capable of reacting
 20 through self condensation or with carboxylic acid by condensation or ring opening; and

(c) a hydrophobic monomer.

25 2. The composition of Claim 1, wherein the reactive self crosslinking monomer (b) is one in which X is selected from the group consisting of glycidyl, N-methylol, N-methylol alkyl ether having alkyl groups of from one to four carbon atoms, blocked isocyanate, N-methylol urethane
 30 or aziridine.

3. The composition of Claim 1, wherein the hydrophobic monomer (c) is selected from the group consisting of alkyl esters of acrylic or methacrylic acid wherein said alkyl
 35 groups have from one to eight carbon atoms.

- 20 -

4. The composition of Claim 3, wherein said alkyl groups have from one to four carbon atoms.

5. The composition of Claim 1, wherein said hydrophobic monomer is present in sufficient quantity to result in 5 to about 70 percent by weight of hydrophobic monomer in said copolymer.

6. The composition of Claim 5, wherein said hydrophobic monomer is present in sufficient quantity to result in 10 to 35 percent by weight of hydrophobic monomer in said copolymer.

7. The composition of Claim 1, wherein said poly(electrolyte) monomer (a) is present in sufficient quantity to result in a copolymer containing about 20 to about 82 percent by weight of carboxylic groups.

8. The composition of Claim 7, wherein said poly(electrolyte) monomer (a) is present in sufficient quantity to result in a copolymer containing about 35 to about 75 percent by weight of carboxylic groups.

9. The composition of Claim 1 in the form of an emulsion wherein said copolymer is in the dispersed phase.

10. The composition of Claim 1 in the form of a viscous solution having a pH of at least 5.

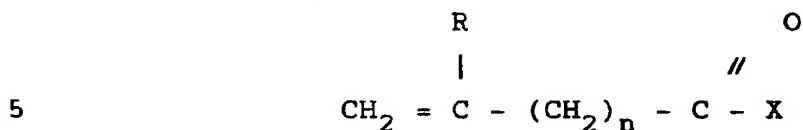
11. An absorbent material comprising poly(electrolyte) self crosslinking copolymer; said copolymer comprising the product of the copolymerization of:

(a) a carboxylic polyelectrolyte monomer;

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(b) a reactive self crosslinking monomer having the formula:



wherein R is selected from the groups consisting of H or CH₃, n is selected from the group consisting of zero or the integers 1 to 5, and X is selected to be a group capable of reacting through self condensation or with carboxylic acid by condensation or ring opening; and

(c) a hydrophobic monomer.

12. The absorbent material of Claim 11, wherein the reactive self crosslinking monomer (b) is one in which X is selected from the group consisting of glycidyl, N-methylol, N-methylol alkyl ether having alkyl groups of from one to four carbon atoms, blocked isocyanate, N-methylol urethane or aziridine.

13. The absorbent material of Claim 11, wherein the hydrophobic monomer (c) is selected from the group consisting of alkyl esters of acrylic or methacrylic acid wherein said alkyl groups have from one to eight carbon atoms.

14. The absorbent material of Claim 13, wherein said alkyl groups have from one to four carbon atoms.

15. The absorbent material of Claim 11, wherein said hydrophobic monomer is present in sufficient quantity to result in 5 to about 70 percent by weight of hydrophobic monomer in said copolymer.

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16. The absorbent material of Claim 15, wherein said hydrophobic monomer is present in sufficient quantity to result in 10 to 35 percent by weight of hydrophobic monomer in said copolymer.

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17. The absorbent material of Claim 11, wherein said poly(electrolite) monomer (a) is present in sufficient quantity to result in a copolymer containing about 20 to about 82 percent by weight of carboxylic groups.

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18. The absorbent material of Claim 17, wherein said poly(electrolyte) monomer (a) is present in sufficient quantity to result in a copolymer containing about 35 to about 75 percent by weight of carboxylic groups.

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19. The absorbent material of Claim 11 in the form of a powder.

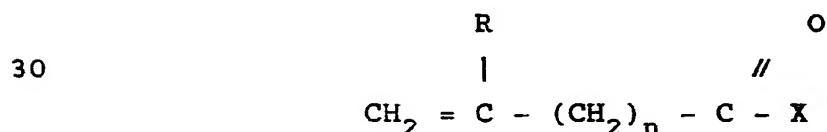
20. The absorbent material of Claim 11 in the form of a film.

20

21. A process for producing absorbent materials comprising:

25 (a) a carboxylic polyelectrolyte monomer;

(b) a reactive self crosslinking monomer having the formula:



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wherein R is selected from the groups consisting of H or CH₃, n is selected from the group consisting of zero or the integers 1 to 5, and X is selected to be a group capable of reacting

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through self condensation or with carboxylic acid
by condensation or ring opening; and

(c) a hydrophobic monomer.

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copolymerizing said monomers to form a poly(electrolyte)
self crosslinking copolymer dispersed in water; and drying
said dispersion to produce absorbent material comprising
crosslinked poly(electrolyte) copolymers.

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22. The process of Claim 21, wherein the reactive self
crosslinking monomer, (b), is one in which X is selected
from the group consisting of glycidyl, N-methylol,
N-methylol alkyl ether having alkyl groups of from one to
15 four carbon atoms, blocked isocyanate, N-methylol
urethane, or aziridine.

23. The process of Claim 21, wherein the hydrophobic
monomer (c) is selected from the group consisting of alkyl
20 esters of acrylic or methacrylic acid wherein said alkyl
groups have from one to eight carbon atoms.

24. The process of Claim 23, wherein said alkyl groups
have from one to four carbon atoms.

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25. The process of Claim 21, wherein said hydrophobic
monomer is present in sufficient quantity to result in 5
to about 70 percent by weight of hydrophobic monomer in
said poly(electrolyte) self crosslinking copolymer.

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26. The process of Claim 25, wherein said hydrophobic
monomer is present in sufficient quantity to result in 10
to 35 percent by weight of hydrophobic monomer in said
copolymer.

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27. The process of Claim 21, wherein said poly(electrolyte) monomer (a) is present in sufficient quantity to result in a copolymer containing about 20 to about 82 percent by weight of carboxylic groups.

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28. The process of Claim 27, wherein said poly(electrolyte) monomer (a) is present in sufficient quantity to result in a copolymer containing about 35 to about 75 percent by weight of carboxylic groups.

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29. The process of Claim 21, wherein the copolymerization is catalyzed by free radical generating compounds.

30. The process of Claim 21, wherein the copolymerization in initiated by oxidation/reduction initiators selected from the group consisting of persulfate, bisulfate and ferrous ions.

31. The process of Claim 21, wherein the shelf life of the dispersion is extended by neutralizing the dispersion to a pH of about 5 or greater.

32. The process of Claim 21, wherein said absorbent material is incorporated into a substrate by applying said dispersion to said substrate and then drying to produce crosslinked poly(electrolyte) in situ.

33. The process of Claim 32, wherein said substrate is a fibrous web.

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34. The process of Claim 32, wherein said substrate is a cellular polymer.

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